

New linker systems for superior immobilized catalysts†

Björn Beele,^{ab} Johannes Guenther,^a Melanie Perera,^a Michaela Stach,^a
Thomas Oeser^b and Janet Blümel^{*a}

Received (in Gainesville, FL, USA) 23rd June 2010, Accepted 27th July 2010

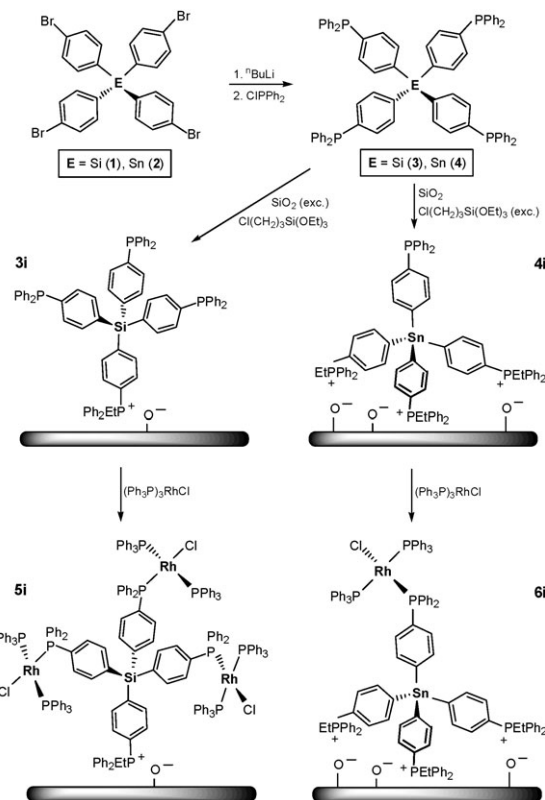
DOI: 10.1039/c0nj00482k

The rigid linkers $E(p\text{-C}_6\text{H}_4\text{PPh}_2)_4$ ($E = \text{Si}, \text{Sn}$) have been synthesized from the corresponding precursors $E(p\text{-C}_6\text{H}_4\text{Br})_4$ and immobilized by forming one ($E = \text{Si}$) or three ($E = \text{Sn}$) phosphonium groups that are attached to the silica support by electrostatic interactions. The remaining unquaternized phosphines are coordinated to Wilkinson-type Rh complexes. The catalyst immobilized by the rigid scaffold with three surface-bound phosphonium groups ($E = \text{Sn}$) displays superior catalytic activity and lifetime with respect to the hydrogenation of dodecene. This catalyst can be recycled 30 times in a batchwise manner under standardized conditions.

Immobilized species are important in such diverse areas as combinatorial chemistry,¹ solid-phase synthesis,² chromatography,³ and catalysis.^{4,5} Immobilized catalysts^{4,5} are of immense academic and industrial interest because they can combine the advantages of homogeneous with those of heterogeneous catalysts. Immobilized catalysts are highly active and selective, while they can easily be separated from the reaction mixture and reused for example in a batchwise manner.^{4,5} The most favorable support materials are silica and neutral alumina,^{3,6} and bifunctional phosphines, such as the chelate phosphine $\text{Ph}_2\text{P}(\text{CH}_2)_3\text{Ph}(\text{CH}_2)_3\text{Si}(\text{OEt})_3$, are popular linkers.^{6,7} The latter have led to very successful immobilized nickel⁸ and rhodium⁹ catalysts for the cyclotrimerization of acetylenes and olefin hydrogenation, respectively.

However, as demonstrated in previous work, when flexible phosphine linkers with high surface coverages are used catalyst deactivation by dimerization can occur.⁹ The catalysts can be “diluted” on the surface to prevent dimerization,⁹ but increasing the amount of bulk material in this way is not desirable from an industrial point of view. Additionally, deactivation of the catalysts or linker destruction by reaction with the oxide surface¹⁰ remains possible.

Here, we present a new rigid linker scaffold that allows the clean immobilization of homogeneous catalysts on oxide supports. This linker scaffold prevents catalyst dimerization or contact with the support surface and leads to unprecedented lifetimes and recyclabilities.



Scheme 1 Syntheses of **3** and **4**, the immobilized linkers **3i** and **4i**, and the catalysts **5i** and **6i**.¹³

The silane $\text{Si}(p\text{-C}_6\text{H}_4\text{Br})_4$ (**1**) (Scheme 1) has been synthesized by the method described by Wuest *et al.*,¹¹ slightly modified as outlined for the stannane $\text{Sn}(p\text{-C}_6\text{H}_4\text{Br})_4$ (**2**).¹²

Although $\text{Si}(p\text{-C}_6\text{H}_4\text{Br})_4$ (**1**) is used extensively as the core building block for constructing networks for various applications,¹¹ its ²⁹Si NMR data and crystal structure have not yet been reported. The value $\delta(^{29}\text{Si}) = -13.76$ ppm for **1** in CDCl_3 solution corresponds very well to the value of SiPh_4 with -13.98 ppm.¹⁴ In the solid-state NMR¹⁵ spectrum of polycrystalline **1** (Fig. 1) without rotation, a wide line signal with the moderate CSA span ($\delta_{11} - \delta_{33}$)^{15a} of 14 ppm and the CSA values $\delta_{11} = -8$, $\delta_{22} = -15$, and $\delta_{33} = -22$ ppm are obtained. Two isotropic lines at $\delta = -14.3$ and -15.0 ppm are visible at 4 kHz rotational speed. This is corroborated by the single crystal X-ray analysis,¹⁶ where two crystallographically independent molecules generate the asymmetric unit (Fig. 1). Due to the tetragonal symmetry one fragment lies upon a fourfold rotation axis, the other one is placed upon a twofold rotation axis.

^a Department of Chemistry, Texas A&M University,
P.O. Box 30012, College Station, Texas 77842-3012, USA.
E-mail: bluemel@tamu.edu

^b Institute of Organic Chemistry, University of Heidelberg,
Im Neuenheimer Feld 270, 69120 Heidelberg, Germany

† Electronic supplementary information (ESI) available: Experimental data of **3**, **4**, and **3i** to **6i**. CCDC 742605 and 742606. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c0nj00482k

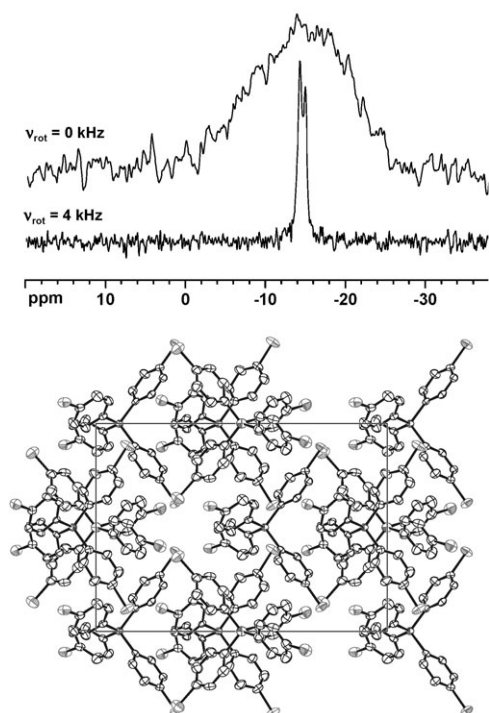


Fig. 1 ^{29}Si wideline (top) and MAS (bottom) NMR spectra of polycrystalline **1**, and unit cell of a single crystal of **1** (view along the crystallographic *a*-axis).¹⁶

As compared to **1**, polycrystalline $\text{Sn}(p\text{-C}_6\text{H}_4\text{Br})_4$ (**2**)¹² shows a rather large CSA span^{15a} of $(\delta_{11} - \delta_{33}) = 40$ ppm in its ^{119}Sn wideline spectrum (Fig. 2), as expected for a heavier nucleus.^{15a} The CSA values $\delta_{11} = -92$, $\delta_{22} = -129$, and $\delta_{33} = -133$ ppm are obtained. Only one isotropic line is visible in the ^{119}Sn MAS spectra of **2** (Fig. 2). This is in correspondence with the single crystal X-ray analysis¹⁷ (Fig. 2), which shows that all Sn nuclei are magnetically equivalent.

For both, the tetraphenylsilane and -stannane systems, the intramolecular distance between the bound catalytic centers should be large enough to prevent their dimerization, as the average intramolecular $\text{Br} \cdots \text{Br}$ distances of 10.53 Å for **1** and 11.07 Å for **2** show.

The phosphines **3** and **4** have been synthesized in yields of 55 and 79% by Br/Li exchange, followed by reaction with ClPPh_2 (Scheme 1).¹³ An alternative synthesis route *via* Br/Li exchange of $p\text{-BrC}_6\text{H}_4\text{PPh}_2$ and reaction with SiCl_4 increases the yield of **3** to 70%.¹³

The silane **3** has been immobilized selectively *via* one phosphonium group on silica¹⁸ to give **3i** (Scheme 1) according to the procedure explored in general,¹⁹ and communicated for rigid tetraphosphines previously.¹² In the meantime, we found reaction conditions that also allow us to bind the rigid tetraphosphines to silica *via* three phosphonium groups.

4i (Scheme 1) can be generated by adding a 100-fold excess of $\text{Cl}(\text{CH}_2)_3\text{Si}(\text{OEt})_3$ to **4** and SiO_2 and stirring at 90 °C for 5 days. The quantitative ^{31}P MAS spectrum^{12,15d} (Fig. 3), recorded with a long relaxation delay of 10 s, proves the number of binding sites per linker molecule by the intensity ratio of *ca.* 3 : 1 for the phosphonium and the phosphine signals at 23.8 and -5.2 ppm.

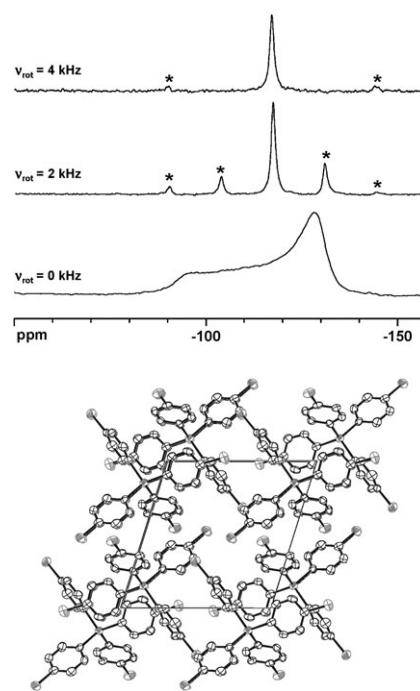


Fig. 2 ^{119}Sn wideline (bottom spectrum) and MAS (top two) NMR spectra of polycrystalline **2** (asterisks denote rotational sidebands), and unit cell of a single crystal of **2** (view along the crystallographic *a*-axis).¹⁷

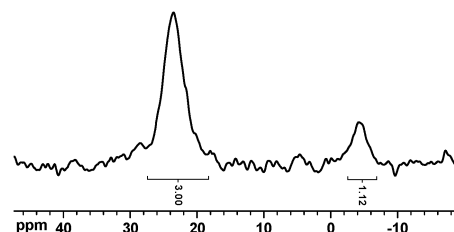


Fig. 3 ^{31}P MAS NMR spectrum of **4i**. $\nu_{\text{rot}} = 4$ kHz.

The Rh catalysts **5i** and **6i** (Scheme 1) have been obtained *via* ligand exchange by stirring **3i** and **4i** with a slight excess of Wilkinson's catalyst, and removing the surplus $\text{ClRh}(\text{PPh}_3)_3$ and PPh_3 by washing with toluene. Both catalysts **5i** and **6i** have been tested with respect to the hydrogenation of 1-dodecene, a catalytic reaction that allows the comparison with ample previous data⁹ under the applied standard conditions described earlier.^{9c}

Since the support material settles down quickly after each catalytic run, **5i** and **6i** could be recycled by decanting the supernatant and washing the solid thoroughly with toluene.

Catalyst **5i** could be recycled 13 times before the consumption of H_2 was no longer 100% within 100 h. This activity and lifetime scenario is comparable to cases where complexes are immobilized *via* flexible chelate linkers such as $\text{Ph}_2\text{P}(\text{CH}_2)_3\text{PPh}(\text{CH}_2)_3\text{Si}(\text{OEt})_3$.⁹ Therefore, if the scaffold is only bound *via* one phosphonium group, the active catalytic centers might still get deactivated by contact with the reactive silica surface.

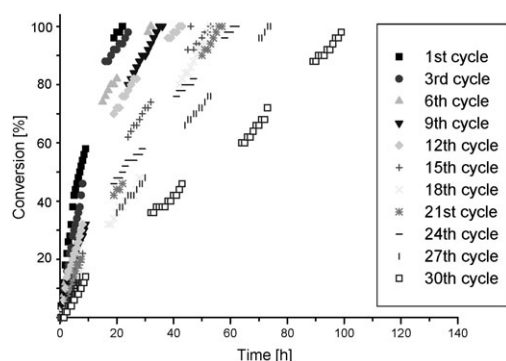


Fig. 4 Batchwise recycling of **6i** after the hydrogenation of 1-dodecene.⁹ Ratio dodecene/Rh: 100/1; concentration of dodecene in toluene: 0.1 mol l⁻¹; pressure 1.1 bar; reaction temperature 25 °C.

In the case of catalyst **6i**, which is bound by a ligand platform with three phosphonium groups, no flexibility should be left. Therefore, no catalyst deactivation by dimerization or contact with the support surface should occur. Accordingly, **6i** showed unprecedented activity and lifetime, and could be recycled for the record number of 30 times under the standard conditions^{9c} (Fig. 4). The ³¹P CP/MAS spectra before and after catalysis show the same signals (see ESI†), in analogy to previously studied successful catalysts with different linker systems.^{9b}

In summary, it has been demonstrated that catalysts with rigid linker scaffolds are easy to synthesize and analyze. These linkers prevent the dimerization of the catalytically active metal centers and their deactivation by contact with the reactive support surface. Especially catalysts surface-bound by tetrahedral systems with three phosphonium binding sites per linker show unprecedented activities and lifetimes.

Acknowledgements

This material is based upon work supported by the National Science Foundation under CHE-0911207. Furthermore we thank The Welch Foundation (A-1706), Texas A&M University, and DFG (SFB 623) for support.

References

- (a) *Handbook of Combinatorial Chemistry*, ed. K. C. Nicolaou, R. Hanco and W. Hartwig, Wiley-VCH, Weinheim, Germany, 2002, vol. 1 and 2; (b) *Combinatorial Chemistry*, ed. W. Bannworth and E. Felder, Wiley-VCH, Weinheim, Germany, 2000.
- (a) P. Seneci, *Solid-Phase Synthesis and Combinatorial Technologies*, John Wiley & Sons, New York, 2000; (b) F. Z. Dörwald, *Organic Synthesis on Solid Phase*, Wiley-VCH, Weinheim, 2000.
- (a) E. F. Vansant, P. VanDer Voort and K. C. Vrancken, *Characterization and Chemical Modification of the Silica Surface*, Elsevier, Amsterdam, 1995; (b) R. P. W. Scott, *Silica Gel and Bonded Phases*, John Wiley and Sons, New York, 1993.
- (a) F. R. Hartley, *Supported Metal Complexes*, D. Reidel Publ. Co., Dordrecht, The Netherlands, 1985; (b) *Chiral Catalyst Immobilization and Recycling*, ed. D. E. DeVos, I. F. J. Vankelecom and P. A. Jacobs, Wiley-VCH, Weinheim, 2000; (c) G. Rothenberg, *Catalysis: Concepts and Green Applications*, Wiley-VCH, Weinheim, 2008.
- J. Blümel, *Coord. Chem. Rev.*, 2008, **252**, 2410.

- (a) C. Merckle and J. Blümel, *Chem. Mater.*, 2001, **13**, 3617; (b) J. Blümel, *J. Am. Chem. Soc.*, 1995, **117**, 2112; (c) K. D. Behringer and J. Blümel, *J. Liq. Chromatogr.*, 1996, **19**, 2753.
- (a) G. Tsiavaliaris, S. Haubrich, C. Merckle and J. Blümel, *Synlett*, 2001, 391; (b) F. Piester, R. Fetouaki, M. Bogza, T. Oeser and J. Blümel, *Chem. Commun.*, 2005, 1481.
- (a) S. Reinhard, P. Soba, F. Rominger and J. Blümel, *Adv. Synth. Catal.*, 2003, **345**, 589; (b) K. D. Behringer and J. Blümel, *Chem. Commun.*, 1996, 653; (c) S. Reinhard, K. D. Behringer and J. Blümel, *New J. Chem.*, 2003, **27**, 776.
- (a) C. Merckle and J. Blümel, *Adv. Synth. Catal.*, 2003, **345**, 584; (b) C. Merckle and J. Blümel, *Top. Catal.*, 2005, **34**, 5; (c) C. Merckle, S. Haubrich and J. Blümel, *J. Organomet. Chem.*, 2001, **627**, 44.
- T. Posset, F. Rominger and J. Blümel, *Chem. Mater.*, 2005, **17**, 586.
- J.-H. Fournier, X. Wang and J. D. Wuest, *Can. J. Chem.*, 2003, **81**, 376.
- Y. Yang, B. Beele and J. Blümel, *J. Am. Chem. Soc.*, 2008, **130**, 3771.
- Analytical data see ESI†.
- R. H. Cragg and R. D. Lane, *J. Organomet. Chem.*, 1984, **277**, 199.
- (a) T. M. Duncan, *A Compilation of Chemical Shift Anisotropies*, Farragut Press, Chicago, IL, 1990; (b) C. A. Fyfe, *Solid-State NMR for Chemists*, C.F.C. Press, Guelph, Canada, 1983; (c) M. J. Duer, *Introduction to Solid-State NMR Spectroscopy*, Blackwell Publishing, Oxford, 2004; (d) S. Reinhard and J. Blümel, *Magn. Reson. Chem.*, 2003, **41**, 406.
- X-Ray data and parameters of **1**: colourless crystal (irregular), dimensions 0.20 × 0.14 × 0.10 mm³, crystal system tetragonal, space group *I*₄, *Z* = 6, *a* = *b* = 13.6296(1) Å, *c* = 18.9768(2) Å, $\alpha = \beta = \gamma = 90^\circ$, *V* = 3525.24(5) Å³, ρ = 1.843 g cm⁻³, *T* = 200(2) K, θ_{\max} = 24.11°, radiation MoK α , λ = 0.71073 Å, 0.3° ω -scans with CCD area detector, covering a whole sphere in reciprocal space, 13 920 reflections measured, 2791 unique (*R*_{int} = 0.0490), 2482 observed (*I* > 2σ(*I*)), intensities were corrected for Lorentz and polarization effects, an empirical absorption correction was applied using SADABS based on the Laue symmetry of the reciprocal space, μ = 6.91 mm⁻¹, *T*_{min} = 0.34, *T*_{max} = 0.55, structure solved by direct methods and refined against *F*² with a Full-matrix least-squares algorithm using the SHELXTL (6.12) software package, 197 parameters refined, hydrogen atoms were treated using appropriate riding models, Flack absolute structure parameter 0.037(16), goodness of fit 1.02 for observed reflections, final residual values *R*₁(*F*) = 0.036, *wR*(*F*²) = 0.090 for observed reflections, residual electron density -0.66 to 0.49 e Å⁻³. CCDC 742605 contains the supplementary crystallographic data for this structure†.
- X-Ray data and parameters of **2**: colourless crystal (irregular), dimensions 0.24 × 0.08 × 0.04 mm³, crystal system triclinic, space group *P*₁, *Z* = 2, *a* = 9.7752(6) Å, *b* = 11.2894(6) Å, *c* = 11.8628(7) Å, α = 70.695(1)°, β = 74.850(1)°, γ = 79.261(2)°, *V* = 1185.4(1) Å³, ρ = 2.081 g cm⁻³, *T* = 200(2) K, θ_{\max} = 22.46°, radiation MoK α , λ = 0.71073 Å, 0.3° ω -scans with CCD area detector, 8197 reflections measured, 3093 unique (*R*_{int} = 0.0493), 2361 observed (*I* > 2σ(*I*)), intensities were corrected for Lorentz and polarization effects, an empirical absorption correction was applied using SADABS based on the Laue symmetry of the reciprocal space, μ = 7.83 mm⁻¹, *T*_{min} = 0.26, *T*_{max} = 0.74, structure solved by direct methods and refined against *F*² with a Full-matrix least-squares algorithm using the SHELXTL (6.12) software package, 262 parameters refined, hydrogen atoms were treated using appropriate riding models, goodness of fit 1.01 for observed reflections, final residual values *R*₁(*F*) = 0.043, *wR*(*F*²) = 0.101 for observed reflections, residual electron density -0.77 to 1.46 e Å⁻³. CCDC 742606 contains the supplementary crystallographic data for this structure†.
- The silica (Merck, 40 Å average pore diameter, 0.063 to 0.2 mm average particle size, specific surface area 750 m² g⁻¹) has been rigorously dried *in vacuo* at 500 °C for 4 days to remove adsorbed water and condense surface silanol groups.
- (a) J. Blümel, *Inorg. Chem.*, 1994, **33**, 5050; (b) J. Sommer, Y. Yang, D. Rambow and J. Blümel, *Inorg. Chem.*, 2004, **43**, 7561.